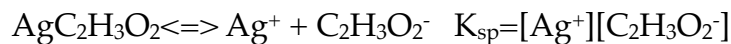
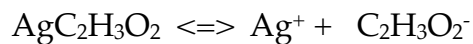


Solubility Product

The solubility of a substance is really an equilibrium situation, most substances will dissolve to some extent, even substances that we have called insoluble. We can define a standard equation for the solubility equilibrium. The solubility product is always defined as a solid breaking apart. The other term that we need to define is solubility. Solubility is the amount of substance that will actually dissolve.



If the solubility of silver acetate is 4.4×10^{-2} what would the K_{sp} be?



	0	0
-4.40E-04	+4.4E-4	+4.4E-4
	4.40E-04	4.40E-04

$$K_{\text{sp}} = [\text{Ag}^+][\text{C}_2\text{H}_3\text{O}_2^-]$$

$$K_{\text{sp}} = [4.4 \times 10^{-4}][4.4 \times 10^{-4}]$$

$$K_{\text{sp}} = 1.9 \times 10^{-7}$$

Ex:

Ex: