

Real Gasses

No gas follows ideal behavior exactly. However, many gases approach ideal behavior at high temperatures and low pressures. We can make some corrections to the ideal gas law that better match how real gases behave.

In 1873 Johannes van der Waals developed an equation for real gases. In 1910 he was awarded a Nobel prize for his work. There are two corrections that van der Waals came up with...

The first correction is to take in to account that real gasses do have volume.

The second correction is to take in to account the attractive forces between gas particles.



Ex: